Unit 2 Test – What you need to know

- Define WHMIS and where/why it is used
- Be able to identify the WHMIS symbols
- Know the difference between physical and chemical properties
- Define:
    - state
    - colour
    - malleability
    - ductility
    - electrical conductivity
    - magnetism
    - lustre
    - melting point
    - boiling point
    - texture
    - combustibility
    - reactivity
    - toxicity
- Know the atomic theory proposed by:
    - John Dalton
    - JJ Thomson
    - Ernest Rutherford
    - Neils Bohr
- Describe an atom including what subatomic particles it contains and where they are located
- Define element
- Write symbols for the 20 elements given
- Know what the Periodic table is and who invented it (including how/when)

- Use the periodic table to identify:
    - Metals
    - Non-metals
    - Metalloids
    - Alkali metals
    - Alkaline earth metals
    - Transition metals
    - Halogens
    - Noble gases

- Know the basic properties of each group above
- Use the periodic table to find an element's atomic number, atomic mass and number of protons, neutrons and electrons in its atoms
- Be able to draw a Bohr-Rutherford diagram for any element in the first 4 rows of the periodic table
- Use the periodic table to determine an element's "period" and "family"
- Be able to determine the number of valence electrons in an atom's valence shell and what this means about how reactive the element is
- Describe a covalent compound and how covalent bonds are made
- Describe an ionic compound and how ionic bonds are made
- Identify compounds as covalent or ionic
- Describe a physical change and provide examples of physical changes
- Describe a chemical change, provide examples of chemical changes and give 5 pieces of evidence of chemical changes

Review questions pg 102 #2, 3, 5, 7, 9, 12, 13, 14, 17, 18, 20, 23, 29, 30, 36